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Medical MCQ 4 All

b) Covalent Character:. Small cation and large anion favors covalency. Order: $\text{LiCl} > \text{NaCl} > \text{KCl} > \text{RbCl} > \text{CsCl}$ & $\text{LiI} > \text{LiBr} > \text{LiCl} > \text{LiF}$. Greater the charge on the cation greater is its polarizing power and hence larger is the covalent character: $\text{Na} + \text{Cl}^- < \text{Mg} + 2 \text{Cl}^- < \text{Al} + 3 \text{Cl}^-$. Greater the charge on the anion, more easily it gets polarized thereby imparting more covalent character to the ...

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COVID-19 testing - Wikipedia

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Pharmacovigilance - A Complete Guide to Pharmacovigilance ...

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Singapore - Wikipedia

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